

**A 3D Polyoxoniobate-based Framework Showing Performance in Dye  
Removal from Aqueous Solution**

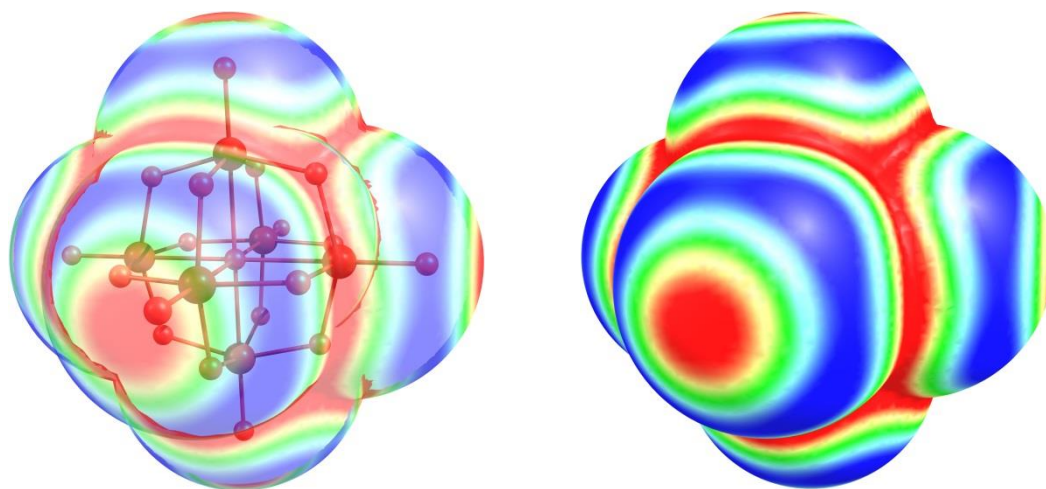
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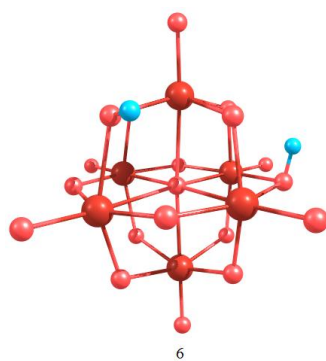
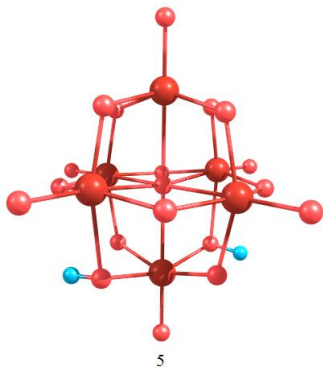
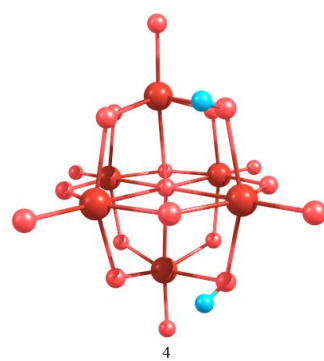
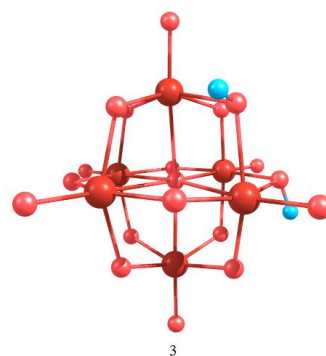
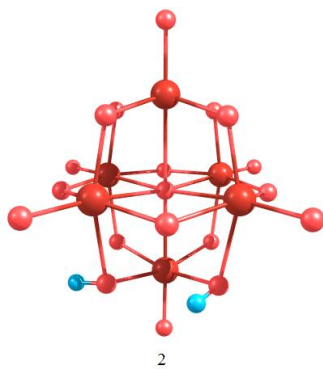
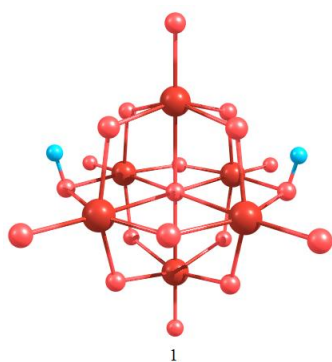
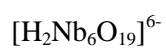
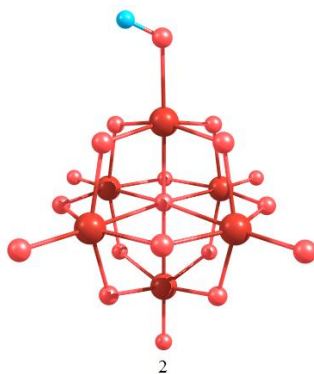
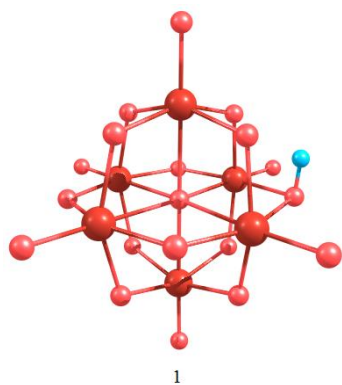
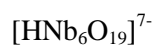
<sup>b</sup> School of Chemical Engineering, University of Tehran, Tehran, Iran

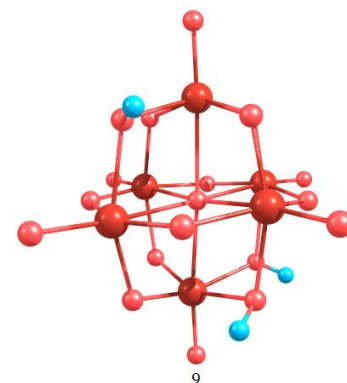
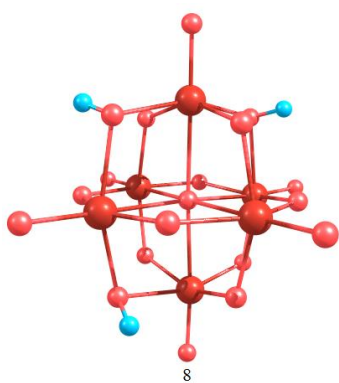
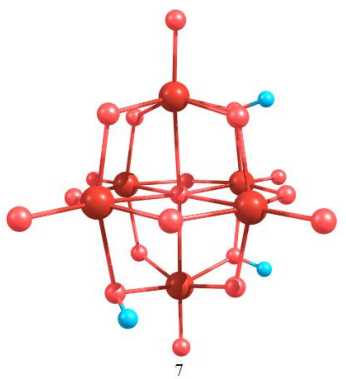
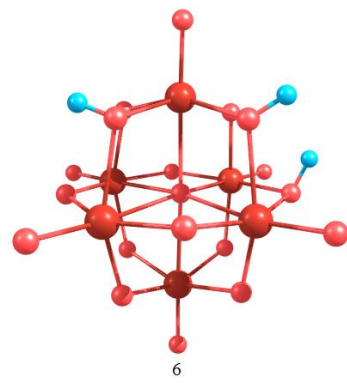
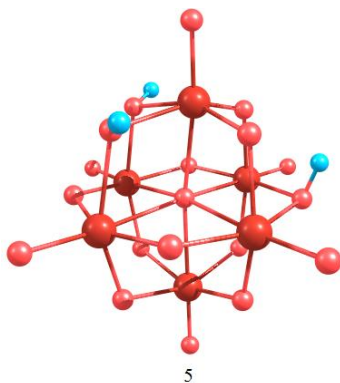
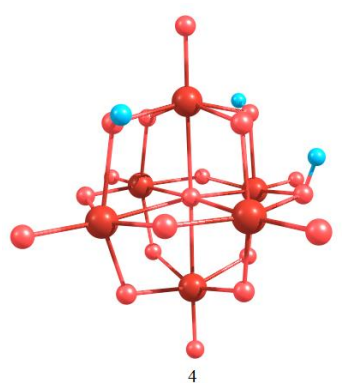
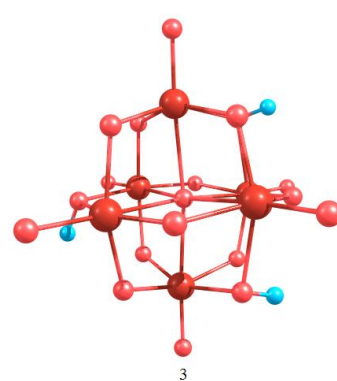
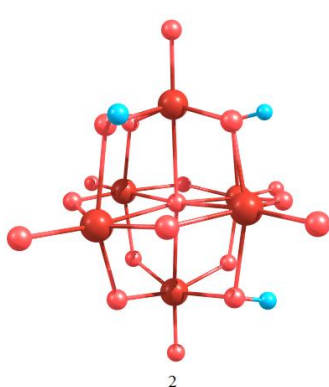
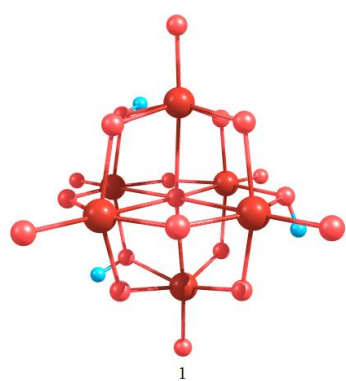
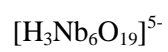
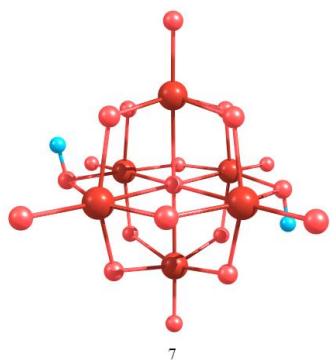
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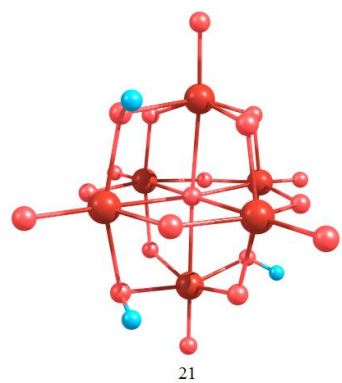
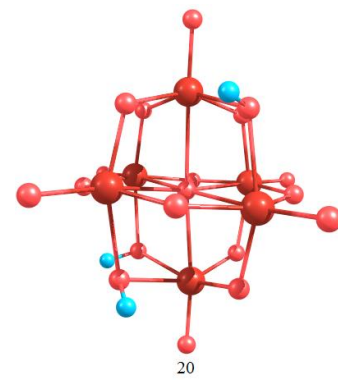
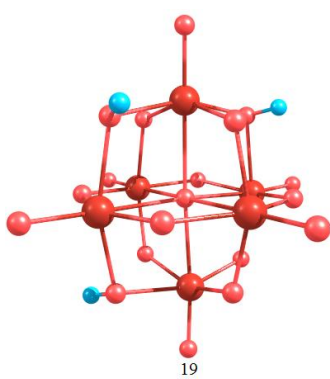
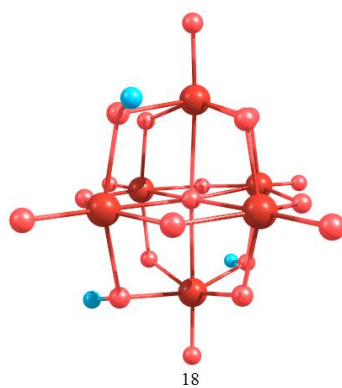
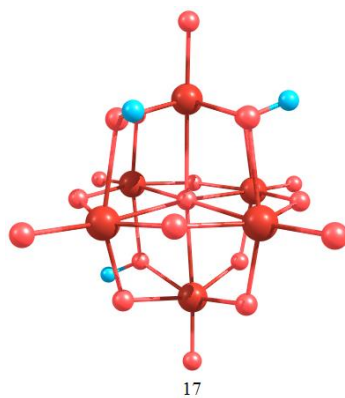
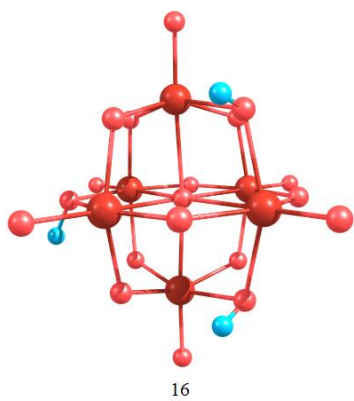
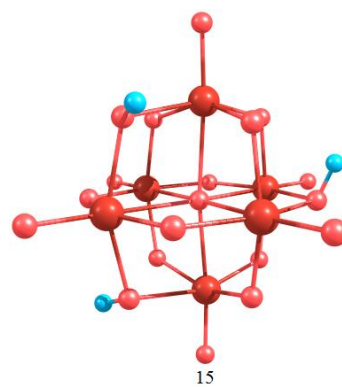
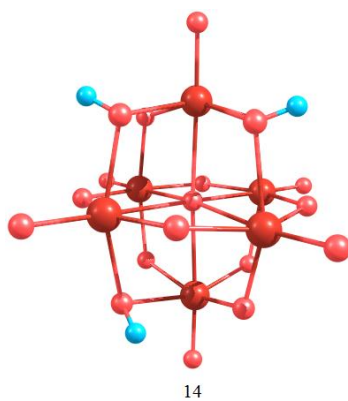
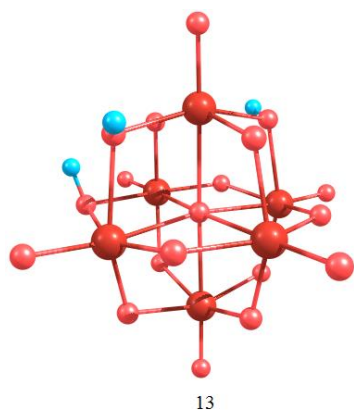
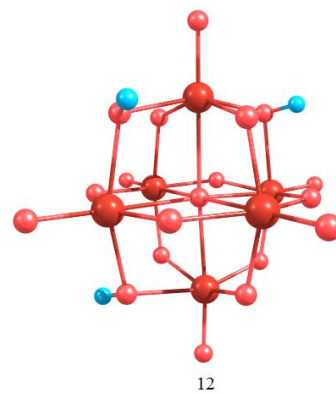
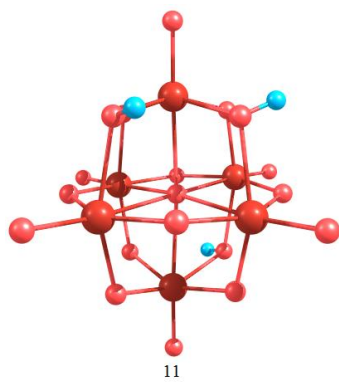
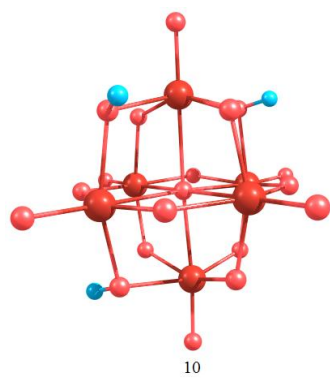
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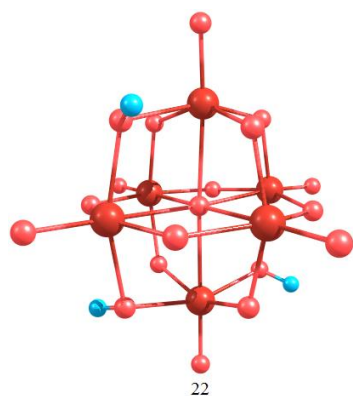


**Figure S1** ESP map for  $\text{Nb}_6\text{O}_{19}^{8-}$

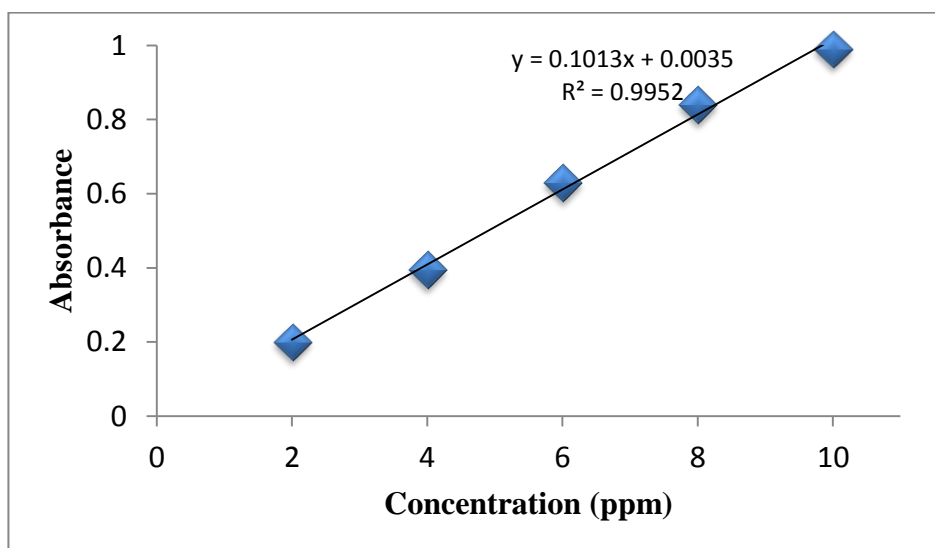




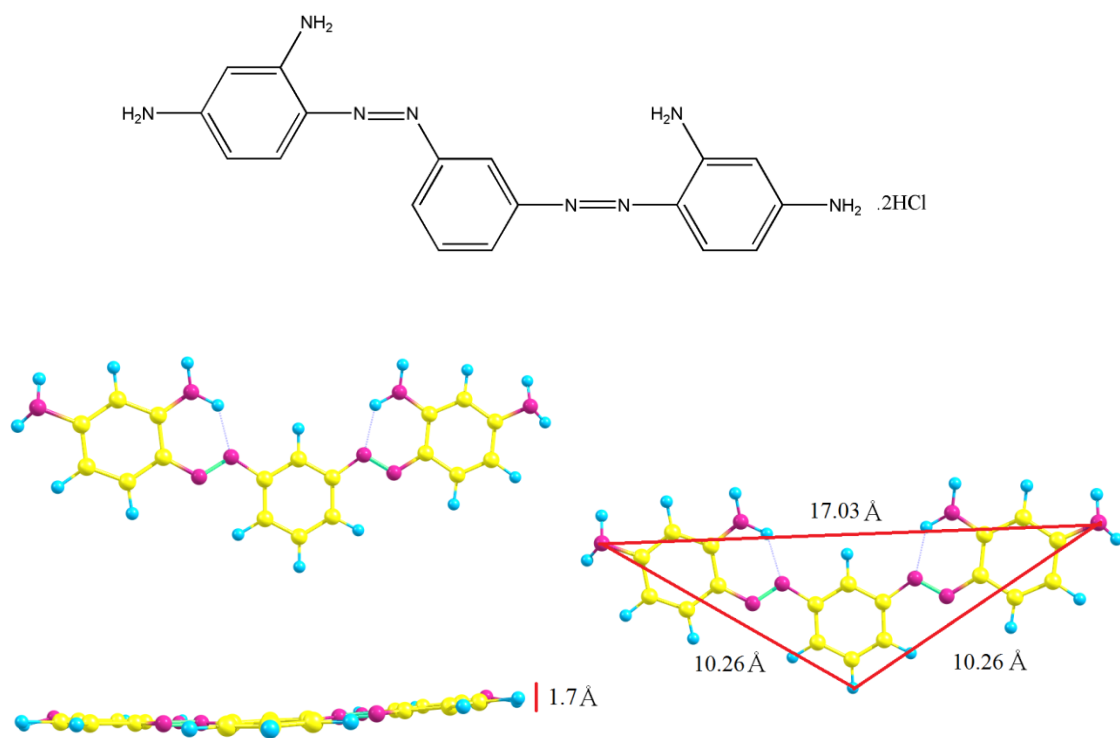




**Figure S2** The lowest energy isomers for each protonation state



**Figure S3** UV-Vis calibration curve



**Figure S4** The chemical structure of Bismarck brown

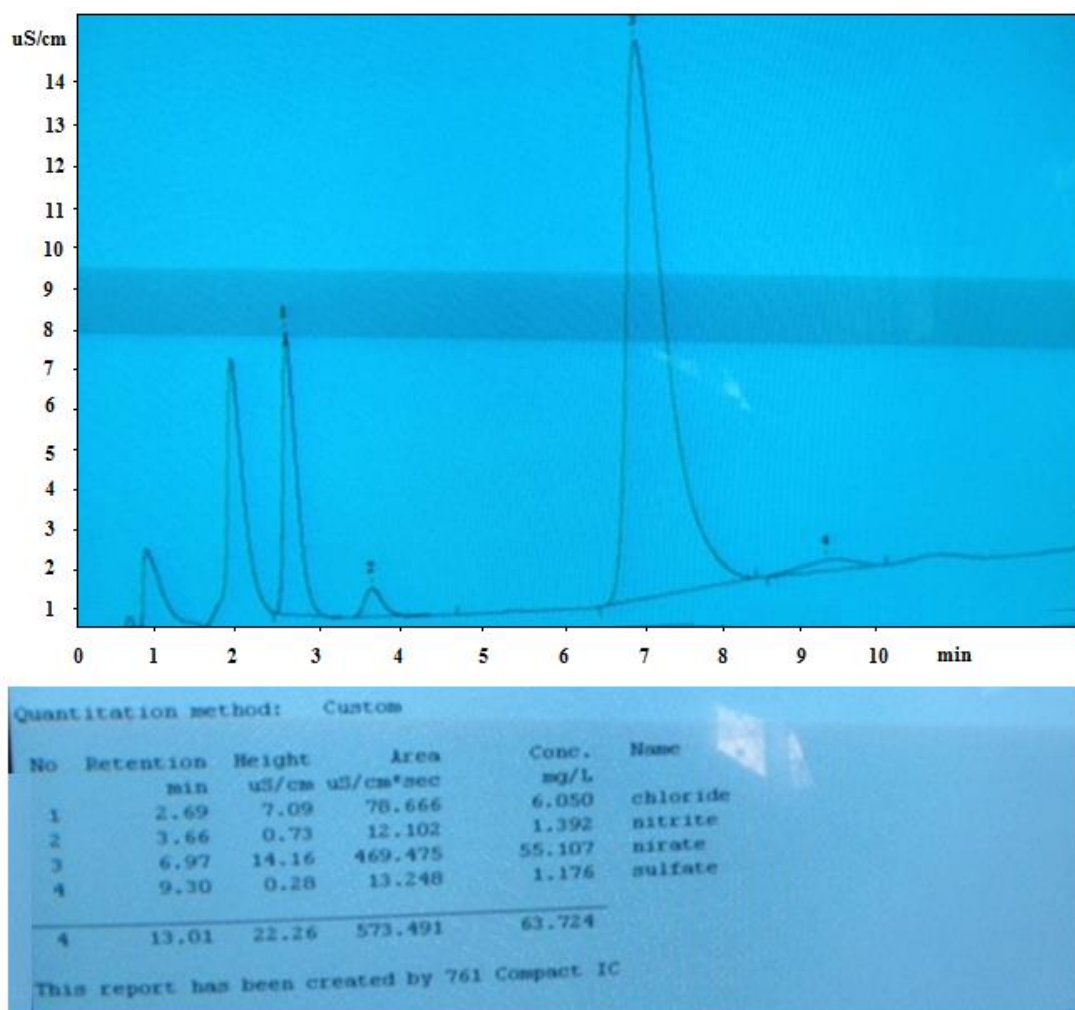


Figure S5 The IC Chromatogram and respected concentrations



**Table S1** Relative energies of hydrogenated hexaniobate

$[\text{HNb}_6\text{O}_{19}]^{7-}$	Relative energy( $\text{kJmol}^{-1}$ )	$[\text{H}_2\text{Nb}_6\text{O}_{19}]^{6-}$	Relative energy( $\text{kJmol}^{-1}$ )	$[\text{H}_3\text{Nb}_6\text{O}_{19}]^{5-}$	Relative energy( $\text{kJmol}^{-1}$ )
$[\text{HNb}_6\text{O}_{19}]^{7-}$ (1)	0.0	$[\text{H}_2\text{Nb}_6\text{O}_{19}]^{6-}$ (1)	13.7	$[\text{H}_3\text{Nb}_6\text{O}_{19}]^{6-}$ (1)	21.1
$[\text{HNb}_6\text{O}_{19}]^{7-}$ (2)	45.0	$[\text{H}_2\text{Nb}_6\text{O}_{19}]^{6-}$ (2)	3.3	$[\text{H}_3\text{Nb}_6\text{O}_{19}]^{6-}$ (2)	20.3
-	-	$[\text{H}_2\text{Nb}_6\text{O}_{19}]^{6-}$ (3)	10.9	$[\text{H}_3\text{Nb}_6\text{O}_{19}]^{6-}$ (3)	0.0
-	-	$[\text{H}_2\text{Nb}_6\text{O}_{19}]^{6-}$ (4)	19.9	$[\text{H}_3\text{Nb}_6\text{O}_{19}]^{6-}$ (4)	21.4
-	-	$[\text{H}_2\text{Nb}_6\text{O}_{19}]^{6-}$ (5)	21.4	$[\text{H}_3\text{Nb}_6\text{O}_{19}]^{6-}$ (5)	4.8
-	-	$[\text{H}_2\text{Nb}_6\text{O}_{19}]^{6-}$ (6)	0.0	$[\text{H}_3\text{Nb}_6\text{O}_{19}]^{6-}$ (6)	11.7
-	-	$[\text{H}_2\text{Nb}_6\text{O}_{19}]^{6-}$ (7)	13.3	$[\text{H}_3\text{Nb}_6\text{O}_{19}]^{6-}$ (7)	49.5
-	-	-	-	$[\text{H}_3\text{Nb}_6\text{O}_{19}]^{6-}$ (8)	50.6
-	-	-	-	$[\text{H}_3\text{Nb}_6\text{O}_{19}]^{6-}$ (9)	11.7
-	-	-	-	$[\text{H}_3\text{Nb}_6\text{O}_{19}]^{6-}$ (10)	49.4
-	-	-	-	$[\text{H}_3\text{Nb}_6\text{O}_{19}]^{6-}$ (11)	12.5
-	-	-	-	$[\text{H}_3\text{Nb}_6\text{O}_{19}]^{6-}$ (12)	49.3
-	-	-	-	$[\text{H}_3\text{Nb}_6\text{O}_{19}]^{6-}$ (13)	21.4
-	-	-	-	$[\text{H}_3\text{Nb}_6\text{O}_{19}]^{6-}$ (14)	27.7
-	-	-	-	$[\text{H}_3\text{Nb}_6\text{O}_{19}]^{6-}$ (15)	22.3
-	-	-	-	$[\text{H}_3\text{Nb}_6\text{O}_{19}]^{6-}$ (16)	20.3
-	-	-	-	$[\text{H}_3\text{Nb}_6\text{O}_{19}]^{6-}$ (17)	12.0
-	-	-	-	$[\text{H}_3\text{Nb}_6\text{O}_{19}]^{6-}$ (18)	49.5
-	-	-	-	$[\text{H}_3\text{Nb}_6\text{O}_{19}]^{6-}$ (19)	49.3
-	-	-	-	$[\text{H}_3\text{Nb}_6\text{O}_{19}]^{6-}$ (20)	11.7
-	-	-	-	$[\text{H}_3\text{Nb}_6\text{O}_{19}]^{6-}$ (21)	49.4
-	-	-	-	$[\text{H}_3\text{Nb}_6\text{O}_{19}]^{6-}$ (22)	50.5