Electrochemical determination of glutathione in hemolyzed erythrocytes

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Multi-walled carbon nanotubes; Glutathione; Erythrocytes; Formazan; Electrochemical determination.

Abstract. The physiological significance of determining glutathione (GSH) and its oxide form is quite clear in terms of its applications in clinical practices such as diagnostic experiments for diabetes, Parkinson's disease, and cancers. Such important determination still needs the development of certain experimental procedures that are easy, fast, and cheap enough to implement. These procedural advantages can be provided through electrochemical methods. Therefore, in this study, on the surface of a Glassy Carbon Electrode (GCE), a composite of functionalized multi-walled carbon nanotubes (MWCNT’s) and formazan was used as a mediator to determine GSH electrochemically. The results indicated that this modified GCE was electrocatalytically very active for glutathione oxidation. Several techniques including Cyclic Voltammetry (CV), scanning electron microscopy, and Differential Pulse Voltammetry (DPV) were applied to characterize the electrode. In addition, such kinetic parameters as the charge transfer rate constant and the transfer coefficient were calculated. In optimized conditions, there was a linear relationship between the DPV peak current of GSH oxidation and GSH concentration in the ranges of 1.0-100.0 and 100.0-800.0 $\mu$M at pH 7.0. As for the detection limit, it was found to be 0.73 $\mu$M. Once put into practice, the devised method proved to be capable of measuring GSH in blood samples.

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1. Introduction

Glutathione (GSH) is known as a biological compound that contains a thiol group. This intracellular thiol compound has a significant role in various biological functions such as redox cycle, electron transfer, protein synthesis, metabolism, and catabolism. Glutathione serves as an enzyme cofactor, disposing poison with high concentrations of heavy metals and setting up protection against damages caused by irradiation. It is also an important factor in increasing the lifespan of Red Blood Cells (RBC), acting as an antioxidant and an inhibitor of free radicals. Glutathione can release free radicals derived from oxygen, which may contribute to the development of various diseases like diabetes [1], Parkinson [2], Alzheimer [3], and some types of cancer [4-6]. The glutathione level in blood can be an index of its level in tissues. Thus, measuring GSH in the blood is a way of identifying disease risks [7].

Glutathione concentration in red blood cells lies within a milli molar range, but within a micro molar range in the plasma. In normal conditions, glutathione concentration in those cells varies in the range of 0.5 to 10 mmol L$^{-1}$ [8,9]. At least, 99% of glutathione exists in

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red blood cells, 16% of which is connected to proteins. Only 0.5% of GSH is detected in the plasma [10]. Therefore, developing a simple, selective and sensitive measuring technique to determine the amount of GSH is important. So far, several chemical methods have been applied to measure GSH: spectrophotometry [11], chromatography [12], analysis of flow injection [13], and enzymatic methods [14]. Most of these methods, however, include difficult steps in sample preparation. Low sensitivity and necessity for derivation are the other disadvantages that account for the reduced use of these methods. In contrast, electrocatalytic methods provide such benefits as simplicity, inexpensiveness, high sensitivity, reproducibility, and easy miniaturization [15].

In this work, following the studies already done by our team to improve the surface of Glassy Carbon Electrodes (GCEs) with different modifiers for the measurement of various biological species and, further, using superior properties of nanoparticles for different methods [16–23], a GCE modified with formazan and multi-walled carbon nanotubes (MWCNTs) is applied to determine GSH concentration. Over recent decades, different electrochemical methods have been used to evaluate GSH in various samples [24–30]. For example, Munteanu et al. developed an electrochemical method to study the effects of the main biological forms of thiols such as homocysteine, cysteine, and GSH (oxidized and reduced forms) on the Cu$^{2+/0}$ system. They could determine GSH at a detection limit of 0.18 µM within the range of 4-7 µM and at a detection limit of 0.05 µM within the range of 0.1-3 µM at pH levels of 1.8 and 4.7, respectively [26]. Ndamanisha et al. used the electrochemical activity of ordered mesoporous carbon (OMC) to measure cysteine (CySH) and glutathione (GSH). They illustrated the great importance of the arranged structure of OMC. It was found that due to its great surface area, appropriate porous structure, and big pore content, OMC can be a promising material for electrocatalysis and devastation of structures that result from the decrement of electrochemical and electrocatalytic properties [29]. Tang et al. reported a procedure for the electrochemical oxidation of glutathione and determined it at an aligned carbon nanotube (CNT) electrode. This electrode showed great electrochemical efficiency for electron transfer reactions. The method was used to determine GSH in the linear range of 0.4–16.4 µM and a detection limit of 0.2 µM [30].

This study develops an inexpensive, simple method to electrochemically determine GSH at a high speed rate, a low detection limit, and a wide linear dynamic range. For this purpose, a GCE was modified with a composite of MWCNTs and (2Z,4E)-3-(3,4-dihydroxyphenyl)-1,5-bis (2,4-dinitrophenyl) formazan (DPDBF). Then, the modified electrode was employed to study the electrocatalytic oxidation of GSH through CV and DPV. It was also of use to assay GSH in human erythrocytes as real samples. These experiments conducted on GSH were based on the electrochemical behavior of thiol compounds, as discussed by Abiman et al. in 2007 [31].

2. Materials and methods

All the solutions were provided freshly with double distilled water. GSH was purchased from Sigma-Aldrich (Taufkirchen, Germany), MWCNTs from Nanostarche Co., and sulfuric acid, nitric acid, Na$_2$H$_2$EDTA, sulfosalicylic acid, phosphoric acid, and ammonia from Merck. Moreover, DPDBF, as a modifier, was synthesized by the following method.

2.1. Procedure for the synthesis of (2Z,4E)-3-(3,4-dihydroxyphenyl)-1,5-bis (2,4-dinitrophenyl) formazan (DPDBF)

A mixture of (2,4-dinitrophenyl)hydrazine (1.0 mmol) and 3,4-dihydroxy benzaldehyde (1.0 mmol, 138 mg) was stirred at room temperature in 5 ml of ethanol for a certain period of time. This occurred with five drops of glacial acetic acid added to it as a catalyst. After completing the conversion, a red solid was obtained. Then, it was filtered and washed with cold ethanol. The resulting 3,4-dihydroxybenzaldehyde-2,4-dinitrophenylhydrazone (318 mg, 1 mmol) together with 500 mg of sodium acetate (solution A) was dissolved in hot methanol (15 ml) under stirring in the reflux conditions and kept at room temperature for the subsequent coupling reaction. Then, 2,4-dinitroaniline (1 mmol, 183 mg), concentrated hydrochloric acid (1 mL), and ice water were mixed and cooled down to 0°C. To this solution was added a solution of 69 mg of NaNO$_2$ and 10 mL of water dropwise under regular stirring. The new solution obtained (solution B) was kept at temperatures 0–5°C. In the next step, the phenyl hydrazine solution (solution A) was placed in an ice-water bath, and the diazonium chloride (solution B) was added to it dropwise under stirring at 0–5°C. After two hours, the resulting cherry red material was filtered, crystallized from ethanol:water (60:40), and dried in vacuo. Finally, DPDBF was characterized by FT-IR and NMR spectroscopy, as presented in Figures S.1 and S.2 in the Supplementary Material. The FTIR spectra of DPDBF indicate absorption bands at 3459 cm$^{-1}$ and 3348, which are attributed to the OH and NH groups, respectively. Moreover, the appearance of the absorption bands at 1616 cm$^{-1}$, 1574 cm$^{-1}$, and 1493 cm$^{-1}$ indicates the existence of aromatic C=C bands, -C=N-hydrazine, and -N=N-azo bands in combination. The absorption bands at 1546 cm$^{-1}$ and 1362 cm$^{-1}$ are the reason for the presence of NO$_2$ group (Figure S.1 in Supplementary Material).
The data obtained from the NMR spectra also show the successful construction of DPBDF (Figure S.2 in Supplementary Material).

2.2. Preparation of functionalized MWCNT and a composite of functionalized MWCNT/DPBDF

The synthesis of carboxylated MWCNT was done in accordance with the reference method, with certain minor modifications made to it [32]. Briefly, 1.0 g of MWCNT was mixed with 75.0 mL of sulfuric acid and 25.0 mL of nitric acid. The mixture was refluxed at 50°C for as long as 24 hours, then cooled down to room temperature. What resulted was a stable black dispersion that was later neutralized with diluted ammonia, filtered, and dried at 85°C for four hours. To prepare an MWCNT/DPBDF composite (C-MWCNT/DPBDF), 10.0 mg of the dried functionalized MWCNT sample was added to 10.0 mL of dimethylformamide (DMF) under ultrasonication for 1 hour. Then, 15 mg of DPBDF was added to it, and the mixture was sonicated for two minutes. The solution was slowly stirred for 36 hours, followed by centrifugation at 6000 rpm for six minutes. The obtained solid residue was dried at 40°C.

2.3. Fabrication of the modified electrode

A GCE polished with alumina slurry was washed with distilled water and dried up in the air. A suspension of C-MWCNT/DPBDF diluted in a solution of ethanol and distilled water (at a volume ratio of 1:1) was dropped on the surface of the GCE. Then, it was left to air-dried. For the optimization purpose, a solution containing 2.0 mg of the modifier (C-MWCNT/DPBDF) was prepared and cast on the surface of the electrode. The effect of the modifier was investigated in various volumes in the range of 1.5-5.0 μL. The optimized amount selected for the experiments was 3.0 μL.

3. Results and discussion

3.1. Characterization of the prepared functionalized MWCNT and C-MWCNT/DPBDF

The functionalized MWCNT and C-MWCNT/DPBDF were characterized by SEM, and the images are provided in Figure 1(a) and (b). Figure 1(a) displays the MWCNTs with approximately 700 nm in length and 47 nm in the outer diameter. CNTs with strong oxidation agents, such as nitric acid and sulfuric acid, can afford the intense etching of their graphitic surface. This results in producing shorter tubes with a high number of erratic sites [33]. Figure 1(b) displays the SEM image of C-MWCNT/DPBDF. It clearly shows the good connection of DPBDF to MWCNTs. This connection increases the surface area and the catalytic effect of the material. It also enhances the uniformity of the surface. As can be seen in the figure, DPBDF is scattered on the surface of the MWCNTs. The results suggest the successful preparation of these compounds.

3.2. Electrochemical characterization of the modified electrode

Kinetic parameters could be calculated by CV. This method was applied to characterize the electrode already modified. The CV tests were done in a phosphate buffer solution at pH 7.0 and a scan rate of 20 mV s⁻¹. Figure 2 shows the CVs of the bare GCE and C-MWCNT/DPBDF/GCE in a phosphate buffer solution. As observed in the curve a, there is no peak in the CV of the bare GCE, while, in the curve b, the electrode modified with C-MWCNT/DPBDF displays reproducible anodic and cathodic peaks with Epa and Epc at 0.281 and 0.173 V, respectively, vs. Ag/AgCl (3.0 M KCl). The inset in Figure 2 displays the CVs of the bare GCE (grey curve) and DPBDF/GCE (red curve) in a phosphate buffer solution. As expected, the DPBDF modified electrode exhibits a redox peak current. The peak of C-MWCNT/DPBDF/GCE has a larger current than the peak DPBDF/GCE, indicating the MWCNT effect. Based on the references presented, it is shown that the DPBDF redox couple possesses a quasi-reversible behavior [34,35].
The electrochemical behavior of C-MWCNT/DPBDF-GCE was investigated at different pH values through CV (Figure 3). Variations of the anodic and cathodic peak potentials, as a function of pH, are shown in Figure 3(a). In both cases, the plots appear as straight-line curves with slopes of about 60.0 mV/pH. This implies that the system follows the Nernst equation, representing the reaction for the transfer of a pair of electrons and a pair of protons. The plots of the cathodic and anodic peak currents (I_p) linearly depend on v in the range of 10-700 mV s^{-1} (Figure 3(b)). This indicates that the redox process occurs in a diffusion-independent surface-confined manner. The transfer coefficient, α, and the apparent rate constant, k = s, for the surface-confined redox couple were considered in accordance with the Laviron procedure. Figure 3(c) reveals the variation of the peak potentials (E_p) as the function of the scan rate logarithm.

As can be observed, at scan rates above 0.2 V s^{-1}, the plots become linear. With the slopes of the linear sections, it is possible to calculate α_c and α_a (i.e., cathodic and anodic transfer coefficients) as the kinetic parameters. The estimated value of α_c is 0.67. The constant for the rate of electron transfer between the modifier and the GCE can be determined by Eq. (1):

\[
\log k_s = \alpha \log(1 - \alpha) + (1 - \alpha) \log \alpha - \log \frac{RT}{nF\vartheta}
\]

\[
-\alpha(1 - \alpha) nF \Delta E_p \frac{2.3RT}{nF\vartheta}
\]

where n stands for the number of the electrons involved in the redox reaction of the modifier, \vartheta represents the potential scan rate, and other signs convey their conventional meanings. Through Eq. (1), the value of \(k_s\) was found to be \(5.0 \times 10^{-3} \text{ s}^{-1}\).

### 3.3. Electrochemical characterization of the modified electrode (C-MWCNT/DPBDF-GCE) for glutathione

A comparison was made for the electrocatalytic oxidation of GSH on the surfaces of a bare GCE and a modified GCE. As observed in Figure 4, the oxidation of GSH on the surface of the modified electrode caused the anodic peak current to rise, but caused the cathodic peak current to decline. This process may be due to the coupling of a chemical reaction and an electron transfer reaction.

The impact of the scan rate upon the electrochemical oxidation of GSH on the surface of the modified electrode was studied through CV (Figure 5(a)). Figure 5(b) shows a plot peak current versus \(v^{1/2}\). It can be understood that the current was diffusion-controlled. The peak potential of GSH oxidation shifted to more positive values with a rise in the scan rate. In this regard, a kinetic limitation could be felt in the electrochemical reaction. As depicted in the figure, due to the rise of the scan rate, the cathodic peak current rose too. This is because, within a short experiment time at a high scan rate, the time would not be enough for a catalytic reaction to occur between oxidized DPBDF and GSH. Accordingly, in this study,
only some of the oxidized DPBDF could be reduced at the reverse scan. Figure 5(c) presents a Tafel plot for the sharply rising section of the current-potential curve.

3.4. Study of interference

The investigation of interferences with electrode responses helps minimize their effects on the preparation of electrodes. This study is partly concerned with how different foreign species affect the determination of 1.0 × 10⁻⁴ M GSH. The maximum concentration of those species was considered as the tolerance limit, leading to a relative error of ±5% in GSH determination. The tolerated concentrations for the foreign materials were 1.0 × 10⁻¹ M for Cl⁻, Na⁺, PO₄³⁻, F⁻, CO₃²⁻, HCO₃⁻ and K⁺, 1.0 × 10⁻² M for glutamic acid and sucrose, 5.0 × 10⁻³ M for ascorbic acid, riboflavin, L-cysteine and glycine, and 2.5 × 10⁻⁵ M for folic acid.

3.5. Calibration curve and detection limit

The dynamic range and the sensitivity of the C-MWCNT/DPBDF-GCE were evaluated in several concentrations of GHS. The findings are presented in Figure 6. The current response declined with a rise in the GSH concentration. These two parameters were linearly correlated within a range of 1.0 μM to 800.0 μM. The peak current plot versus the GSH concentration had two linear sections with slopes of −0.437 and −0.0234 μA μM⁻¹ in the concentration range of 1.0−100.0 μM and 100.0–800.0 μM, respectively. The low sensitivity of the second linear part may have been caused by certain kinetic constraints. A detection limit (3σ) of 0.73 μM was found for GSH.

3.6. GSH determination in real samples

Glutathione was determined successfully in several blood samples using the C-MWCNT/DPBDF-GCE.
Table 1. Recovery values for the voltammetric determination of a GSH standard solution added to three blood samples.

<table>
<thead>
<tr>
<th>Sample</th>
<th>GSH added (µM)</th>
<th>GSH found (µM)</th>
<th>Recovery (%)</th>
<th>RSD</th>
</tr>
</thead>
<tbody>
<tr>
<td>1</td>
<td>0.0</td>
<td>54.146</td>
<td></td>
<td></td>
</tr>
<tr>
<td></td>
<td>2.0</td>
<td>56.126 (±0.361)</td>
<td>99.00</td>
<td>0.643</td>
</tr>
<tr>
<td>2</td>
<td>0.0</td>
<td>53.414</td>
<td></td>
<td></td>
</tr>
<tr>
<td></td>
<td>2.0</td>
<td>55.458 (±0.729)</td>
<td>102.20</td>
<td>1.314</td>
</tr>
<tr>
<td>3</td>
<td>0.0</td>
<td>27.561</td>
<td></td>
<td></td>
</tr>
<tr>
<td></td>
<td>2.0</td>
<td>29.487 (±0.314)</td>
<td>98.32</td>
<td>1.065</td>
</tr>
</tbody>
</table>

Briefly, 5 mL of each blood sample was put in a test tube, and its erythrocytes were separated by eliminating the plasma. For this purpose, the sample was centrifuged for 15 minutes at 3000 rpm and 10⁰C. The plasma that floated on the surface was discarded. What remained were erythrocytes that were then washed with a phosphate buffer solution (pH 7.0) and 0.9% NaCl for three times to eliminate any trace of the plasma left there. The erythrocyte pellets were hemolyzed in a 1.0 mmol L⁻¹ Na₂H₂EDTA solution (1:1, v/v). Then, 5-sulfosalicylic acid (10% (w/v) was added to the prepared hemolyzed erythrocytes at a ratio of 2:1 (v/v). The obtained solution was put under centrifuge at 3000 rpm for 10 minutes. The supernatant material was collected for GSH measurements [36,37]. To check the applicability of this analytical method for real samples, it was used to measure GSH in three blood samples taken from people of different ages (32, 39, and 56 years). After separating the erythrocytes and diluting the solution according to the above-mentioned method, the samples were analyzed to determine GSH. To check against the matrix effect, the glutathione value was determined through a standard addition method. For this purpose, 20.0 µL of 1000.0 µM GSH spiked to 10 mL of each sample and, then, was analyzed through the described procedure. Table 1 presents the results of GSH determination in three different blood samples. Recovery of the empirical results was satisfactorily done for glutathione. The corresponding mean Relative Standard Deviation (RSD) confirmed the reproducibility of this method.

4. Conclusion

In this research, a new electrochemical biosensor was applied to determine GSH in real samples. Formazan was used as a modifier on the surface of a GCE, and that modifier proved to have a superior catalytic effect on GSH oxidation. Through cyclic voltammetry, it was found that GSH oxidation took place at a potential of about +0.3 V. The proposed procedure was shown to be fast, selective, and sensitive. It is also responsible for GSH determination and can be used to analyze real samples. Table 2 shows the results of this method and several other electrochemical methods practiced

Table 2. Comparison of the analytical data for the electrochemical determination of GSH.

<table>
<thead>
<tr>
<th>Electrode</th>
<th>Modifier</th>
<th>pH</th>
<th>Detection limit (µM)</th>
<th>Linear range (µM)</th>
<th>Sensitivity (µA µM⁻¹)</th>
<th>Ref.</th>
</tr>
</thead>
<tbody>
<tr>
<td>GCE</td>
<td>Graphen oxide nanosheets/hybrid of copper-cobalt hexacyano ferrate</td>
<td>3.0</td>
<td>0.25</td>
<td>0.33-55</td>
<td>0.0682</td>
<td>[38]</td>
</tr>
<tr>
<td>Nanocarbon paste electrode</td>
<td>Catechol</td>
<td>7.5</td>
<td>3.2</td>
<td>1-100</td>
<td>0.014</td>
<td>[39]</td>
</tr>
<tr>
<td>GCE</td>
<td>Cobalt phthalocyanine/ MWCNT</td>
<td>7.4</td>
<td>100</td>
<td>500-7000</td>
<td>9.7 µA mM⁻¹</td>
<td>[40]</td>
</tr>
<tr>
<td>GCE</td>
<td>Copper (II) ions</td>
<td>-</td>
<td>0.14</td>
<td>1-12.5</td>
<td>0.1</td>
<td>[41]</td>
</tr>
</tbody>
</table>
for GSH determination. As suggested based on the comparative results, the method proposed in this study is advantageous in certain ways.

Acknowledgements

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Supplementary Material

Supplementary material is available at: http://scientiairanica.sharif.edu/jufile?ar_sfile=130305

References


Table 2. Comparison of the analytical data for the electrochemical determination of GSH (continued).

<table>
<thead>
<tr>
<th>Electrode</th>
<th>Modifier</th>
<th>pH</th>
<th>Detection limit (µM)</th>
<th>Linear range (µM)</th>
<th>Sensitivity (µA µM⁻¹)</th>
<th>Ref.</th>
</tr>
</thead>
<tbody>
<tr>
<td>CPE</td>
<td>Reduced graphene oxide/hydroquinone derivative (DDDC)</td>
<td>7.0</td>
<td>0.02</td>
<td>0.008-100</td>
<td>0.043</td>
<td>[12]</td>
</tr>
<tr>
<td>CPE</td>
<td>Co-based metal-organic coordination polymer (Co-MOCP)</td>
<td>5.5</td>
<td>80</td>
<td>2.5-950</td>
<td>-</td>
<td>[13]</td>
</tr>
<tr>
<td>ITO</td>
<td>Annealed Nickel ion</td>
<td>5.0</td>
<td>5</td>
<td>5-180</td>
<td>66.7 µAc m² ² µM⁻¹</td>
<td>[44]</td>
</tr>
<tr>
<td>Ni-Al layered double hydroxides</td>
<td>Ni-Al LDHs/MWCNTs</td>
<td>-</td>
<td>0.7</td>
<td>1.2-1630</td>
<td>120 µAc m² ² µM⁻¹</td>
<td>[45]</td>
</tr>
<tr>
<td>SPE</td>
<td>CNTs</td>
<td>7.0</td>
<td>3.0</td>
<td>10-100</td>
<td>-</td>
<td>[46]</td>
</tr>
<tr>
<td>GCE</td>
<td>Cu/poly(thionine) (PTH)</td>
<td>8.0</td>
<td>3.5 x 10⁻³</td>
<td>0.01-1000</td>
<td>-</td>
<td>[47]</td>
</tr>
<tr>
<td>GCE</td>
<td>c-MWCNTs/DPBDF</td>
<td>7.0</td>
<td>0.73</td>
<td>1.0-100.0</td>
<td>0.4371</td>
<td></td>
</tr>
</tbody>
</table>


**Biographies**

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